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CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW

SHORT ANSWER Answer the following questions in the

space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient x in this equation? 40.07

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g/mol b. What is the molar mass of C_3H_4 ? 2 mol O
2:1 mol H 2O c. What is the mole ratio ...

~~mc06se-cFMsr-i-vi~~

Chapter 9 Review Stoichiometry Section CHAPTER 9
REVIEW Stoichiometry SECTION 3 PROBLEMS Write
the answer on the line to the left. Show all your work
in the space provided. 1. 88% The actual yield of a
reaction is 22 g and the theoretical yield is 25 g.
Calculate the percentage yield. 2. 6.0 mol of N_2 are
mixed with 12.0 mol of H

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SECTION 1 Defining Stoichiometry 368 Chapter 11 •

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~~Chapter 12 Stoichiometry Section Review Answer Key~~
MANUAL Section 11.1 Defining Stoichiometry pages
368-372 Practice Problems pages 371-372 1.
Interpret the following balanced chemical equations
in terms of particles, moles, and mass. Show that the
law of conservation of mass is
StoichiometryStoichiometry - Weebly Chapter 11
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Section 10.1 shows the general equation
stoichiometry steps as measurable property 1 moles 1
moles 2 measurable property 2 When the reactants
and products of a reaction are pure solids and pure
liquids, mass is the conveniently measurable
property, but many chemical changes take place in
either the gas phase or in solution.

~~Chapter 10 Chemical Calculations and Chemical
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A reaction stoichiometry problem in which you are

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given the mass of B and asked to calculate the number of moles of C produced, you are solving a mass-mole problem The reactant that controls the amount of product formed in a chemical reaction is called the

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the answer on the line to the left. Show all your work
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Calculate the percentage yield. 2.

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STOICHIOMETRY The magic of chemistry happens in this unit! We will learn all about moles, a measure of chemical quantity. We will see how stoichiometry can be used to predict yield in a chemical reaction while doing a lot of hands on chemical reactions in the unit!

~~McLaughlin, Kimberly / Stoichiometry~~

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Chemistry Final EXAM Review Chapters 9-16 &
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Review Describe a chemical reaction. Define reactant. Define product. Identify the products and

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Review ... CHAPTER 9 REVIEW Stoichiometry SECTION
3 PROBLEMS Write the answer on the line to the left
Show all your work in the space provided 1 88% The
actual yield of a reaction is 22 g and the theoretical
yield is 25 g Calculate the percentage yield 2 60 mol
of N₂ are mixed with 120 mol of H

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