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Answer: Calculate the number of half-lives;  
 $0.003 \text{ seconds} \times 1 \text{ half-life} = 3 \text{ half-lives}$

0.001 second • After 0 half-lives, 10 g are left. After 1 half-life, 5 g are left. After 2 half-lives, 2.5 g are left.

After 3 half-lives, 1.25 g are left.

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is radioactivity? 2.  
What is half-life? ...  
Use the following  
graph to answer  
questions 7-10. 7.

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How long is a half-life  
for carbon-14? \_\_\_\_\_

8. If only 25% of the  
carbon-14 remains,  
how old is the  
material containing  
the carbon-14? \_\_\_\_\_

9. If a sample  
originally had 120  
atoms of carbon-14,  
how many atoms will  
remain ...

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## HALF-LIFE

### WORKSHEET

### HALF-LIFE

### WORKSHEET Name

Use Reference Table  
on side to assist you  
in answering the  
following questions.

Equations: Yt lifes:

As-81 = 33 seconds

Au-198 = 2.69 days

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## HALF-LIFE PROBLEMS

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If each half life is 5 seconds, then in one minute. (60 seconds) there are 12 half lives. Therefore the answer is: 2) Iodine-131 has a half life of 8 days. What fraction of the original sample would remain at the end of 32 days? Using the same fraction, you must figure out n. If the half life is 8

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60 seconds/ 5

seconds = 12 half

lives. 1 (  $\frac{1}{2}$  (  $\frac{1}{4}$  (  $\frac{1}{8}$  (

$\frac{1}{16}$  (  $\frac{1}{32}$  (  $\frac{1}{64}$  (

$\frac{1}{128}$  (  $\frac{1}{256}$  (  $\frac{1}{512}$

(  $\frac{1}{1024}$  (  $\frac{1}{2048}$  (

$\frac{1}{4096}$  will remain.

Iodine-131 has a half

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With Answers. What  
fraction of the  
original sample  
would remain at the  
end of 32 days? 32  
days / 8 days = 4 half  
lives.  $1 \left( \frac{1}{2} \right) \left( \frac{1}{4} \right) \left( \frac{1}{8} \right) \left( \frac{1}{16} \right)$   
will remain.

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designed for an  
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Problem #3: Os-182  
has a half-life of 21.5  
hours. How many

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grams of a 10.0 gram sample would have decayed after exactly three half-lives?

Solution:  $(1/2)^3 = 0.125$  (the amount remaining after 3 half-lives)  $10.0 \text{ g} \times 0.125 = 1.25 \text{ g}$  remain  $10.0 \text{ g} - 1.25 \text{ g} = 8.75 \text{ g}$

have decayed Note that the length of the half-life played no role in this

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