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proportions. atom. States that matter can be neither destroyed nor created during a chemical reaction. The fact that a chemical compound contains the same elements in the same proportions by mass. If two or more different compounds are composed of the same elements in the same proportions by mass, they are said to be isomorphous.

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conservation of mass- mass is neither
created nor destroyed during ordinary
chemical reactions or physical
changes. e.g. 20 g A + 20 g B 40 g

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AB.

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 2

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SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in molecules containing the

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3 Atoms: The Building Blocks of Matter

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Unformatted text preview: Modern Chemistry Chapter 3 Atoms: The Building Blocks of Matter law of conservation of mass mass is neither created nor destroyed during ordinary chemical reactions or physical changes e.g. $20\text{ g A} + 20\text{ g B} = 40\text{ g AB}$ law of definite proportions- a chemical

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compound contains the same elements in exactly the same proportions by mass regardless of the size of the sample or the source of the compound e.g. If 10 grams of A combine with 20 grams of B to form compound AB, how ...

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CHAPTER 3 TEST Class Atoms: The Building Blocks of Matter MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice best completes the statement or answers the question. The behavior of

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cathode rays in a glass tube containing gas at low pressure led scientists to conclude that the rays were composed of a. energy

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