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oxidation. A process that involves complete or partial loss of electrons or a gain of oxygen. It results in an increase in the oxidation number of an atom. To balance a redox reaction using half-reactions, write separate half-reactions for the oxidation and the reduction.

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An oxidation-reduction reaction is any chemical reaction in which the oxidation number of a molecule, atom, or ion changes by gaining or losing an electron. Redox reactions are common and vital to some of the basic functions of life, including photosynthesis , respiration , combustion , and corrosion or rusting.

11.4: Oxidation States & Redox Reactions - Chemistry ...

The element undergoing oxidation is the one which assumes a more positive oxidation state when going from reactants and products. The element undergoing reduction is the one which assumes a more negative oxidation state when going from reactants and products. Notes.

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